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1-2. Thermodynamics and Heat Transfer. 1-1C Thermodynamics deals with the amount of heat transfer as a system undergoes a process from one equilibrium state to another.

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1-9C Energy can be transferred by heat, work, and mass. An energy transfer is heat transfer when its driving force is temperature difference. 1-10C Thermal energy is the sensible and latent forms of internal energy, and it is referred to as heat in daily life. 1-11C For the constant pressure case. This is because the heat transfer to an ideal ...

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Table Q1.1 Mean heat capacity of the solution Heat of crystallisation Overall coefficient of heat transfer Latent heat of vaporization for water at 2 kPa Boiling point of water at 2 kPa 3.2 kJ kgK⁻¹ 146.5 kJ kg⁻¹ 140 W m² K⁻¹ 2.5 MJ kg⁻¹ 291 K (a) Calculate the required feeding rate of initial solution (in kg s⁻¹). (8 marks) (b) Calculate the length of crystalliser required under each heat ...

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2.107 Mass of H₂SO₄ solution = 1.3028 g/mL x 40.00 mL = 52.112 g Total mass of Zn and H₂SO₄ solution before reaction = 9.520 g + 52.112 g = 61.632 g Total mass of solution after the

reaction = 61.338 g Because of the conservation of mass, the difference between the two masses is the mass of H₂ produced. H₂ mass = 61.632 g S 61.338 g = 0.294 g ...

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The heat power is calculated using equation , (5) $\dot{q}_{th} = C_p m \dot{q} (T_s - T_0)$ where \dot{q}_{th} is the vector of heat power (W) consumed or supplied at each node; the subscript th represents thermal; C_p is the specific heat of water (J kg⁻¹ °C⁻¹), $C_p = 4.182 \times 10^{-3}$ MJ kg⁻¹ °C⁻¹; and $m \dot{q}$ is the vector of the mass flow rate (kg/s) through each node injected from a supply or ...

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Energy transfer can be considered for the special case of systems which are closed to transfers of matter. The portion of the energy which is transferred by conservative forces over a distance is measured as the work the source system does on the receiving system. The portion of the energy which does not do work during the transfer is called heat.

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